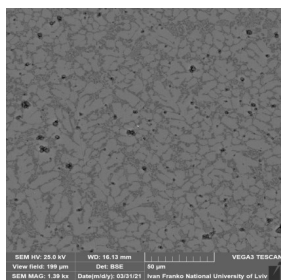


NEW TERNARY COMPOUND  $\text{GdMn}_{1-x}\text{Zn}_{1+x}$ ,  $-0.075 \leq x \leq 0.075$ 

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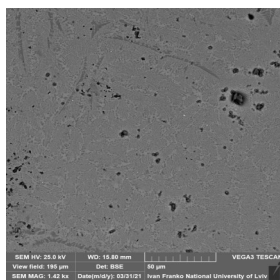
Systematic investigation of the ternary system Gd–Mn–Zn is driven by a practical interest in this type of systems due to their possible application as hydrogen sorption materials and electrode materials for NiMH-batteries. Previous studies revealed the formation of few solid solutions on the basis of binary phases with homogeneity ranges not exceeding 5 at. % and three new ternary compounds –  $\text{Gd}_2\text{Mn}_3\text{Zn}_{14}$  with  $\text{Th}_2\text{Zn}_{17}$ -type structure (space group  $R\bar{3}m$ ) [1],  $\text{GdMn}_{1.075-0.925}\text{Zn}_{0.925-1.075}$  and  $\sim\text{Gd}_{23(1)}\text{Mn}_{46(1)}\text{Zn}_{31(1)}$  with unknown crystal structures and slight homogeneity ranges. Solid solutions on the basis of  $\text{GdMn}_2$  demonstrated hydrogen sorption ability [2].

Samples of alloys were synthesized by arc melting of pure metals (5 wt. % excess of Mn and Zn) under a purified argon atmosphere. The alloys were remelted one time for better homogenization and annealed at  $500^\circ\text{C}$  for two months. X-ray phase analysis and energy dispersive X-ray spectroscopy (TESCAN VEGA3 LMU microscope, EDX-analyzer with X-Max<sup>N</sup>20 detector) showed the formation of a new compound  $\text{GdZn}_{1-x}\text{Mn}_x$  (Fig.). For instance, the alloy  $\text{Gd}_{33.3}\text{Mn}_{33.3}\text{Zn}_{33.3}$  contained two phases – a solid solution  $\text{GdZn}_{1-x}\text{Mn}_x$  and a ternary compound  $\text{GdMn}_{1-x}\text{Zn}_{1+x}$ . The alloy  $\text{Gd}_{33}\text{Mn}_{27}\text{Zn}_{40}$  contained three phases – a solid solution  $\text{GdZn}_{1-x}\text{Mn}_x$ , a ternary compound  $\text{GdMn}_{1-x}\text{Zn}_{1+x}$  and an unknown ternary phase of preliminary composition  $\sim\text{Gd}_{23(1)}\text{Mn}_{46(1)}\text{Zn}_{31(1)}$ . The refinement of the crystal structure of  $\text{GdMn}_{1-x}\text{Zn}_{1+x}$ ,  $-0.075 \leq x \leq 0.075$  from powder data by least square procedure proved that it crystallizes in a hexagonal  $\text{CaIn}_2$ -type structure (space group  $P6_3/mmc$ ,  $a = 4.192(2)\text{ \AA}$ ,  $c = 7.08(1)\text{ \AA}$ ,  $V = 107.8(1)\text{ \AA}^3$ ). Investigation of the hydrogen sorption and electrochemical properties of this compound will be a topic of our further studies.



Light grey phase –  $\text{Gd}_{48.1(5)}\text{Mn}_{5.1(8)}\text{Zn}_{46.8(7)}$ ;  
 Grey phase –  $\text{Gd}_{34.1(3)}\text{Mn}_{32.8(6)}\text{Zn}_{33.1(9)}$

a



Light grey phase –  $\text{Gd}_{48.1(3)}\text{Mn}_{4.0(7)}\text{Zn}_{47.9(7)}$ ;  
 Grey phase –  $\text{Gd}_{34.4(4)}\text{Mn}_{30.9(7)}\text{Zn}_{34.7(7)}$ ;  
 Dark grey phase –  $\text{Gd}_{23(1)}\text{Mn}_{46(1)}\text{Zn}_{31(1)}$

b

Fig. SEM-images (BSE-mode) of  $\text{Gd}_{33.3}\text{Mn}_{33.3}\text{Zn}_{33.3}$  (a) and  $\text{Gd}_{33}\text{Mn}_{27}\text{Zn}_{40}$  (b) alloys

[1] N. Chorna, O. Zelinska, G. Dmytriv, V. Pavlyuk, A. Zelinskiy, V. Kordan, A. Mar. *Coll. Abs. XIV Int. Conf. Cryst. Chem. Intermet. Compd.* Lviv, September 22-26, 2019. P. 119.

[2] N.O. Chorna, V.M. Kordan, A.M. Mykhailevych, O.Ya. Zelinska, A.V. Zelinskiy, K. Kluziak, R.Ya. Serkiz, V.V. Pavlyuk. *Vopr. Khim. Khim. Tekhnol.* 2 (2021) 139-149.